

Is it possible that quantum mechanics is false/wrong?



Michael Brenner

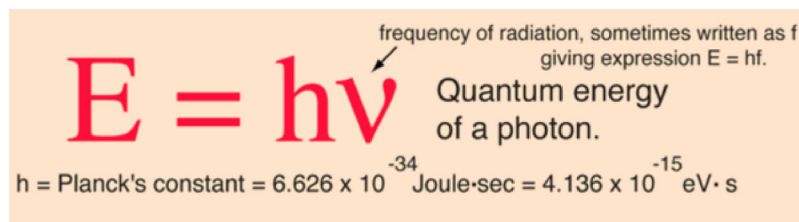
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It is not only possible, it definitely is the case that it is wrong, and it is not that difficult to see either: one doesn't need to have a PhD in theoretical physics, one just needs to be attentive and listen to the language used in describing it, and with language I mean also mathematics. Just force scientists to spell out what they love to write in cryptic math lingo, and it quickly becomes apparent that they either never have done so and are surprised to hear what nonsense they are peddling, or it turns out they actively tried to hide the false and the flawed in cryptic math lingo in order to get a pet theory through publication. Heisenberg was very well aware of that:

"The absurdity of the wave/particle duality of light was not solved but hidden somewhere in the math"

The "Quantum Flaw" has ruined one half of science in the 20th century, the rest was done in by the "Relativity Flaw" which is the interpretation of [c] as "the square root of per inductivity" dressed up as displacement of space in time. $\mu\epsilon=1/c^2$ is the end of physics, the mathematical procedure of inverting and pulling the square root has no representation in nature: $c=\sqrt{1/\mu\epsilon}$

In Quantum we find something similar: the starting point is this equation



This is Planck's ad hoc assumption that radiant energy could only exist in discrete quanta. Interesting that [h] stands for German word for "helper" = "Helfer" because Planck had no idea what it was, and called it "an act of desperation."

This equation could also be called the "one-second-equation" as it is hard coded into the time frame of one second by the term [v] or [f], and there is where the flaw lies buried. Frequency is "Cycles per second" and that means you have to wait for one second of time for all the cycles to come through. Because the equation has no time variable, you are stuck within the one second frame.....and this should already begin to sound weird to you,

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long.

The next flaw is that this equation is not balanced, and thus cannot represent physical reality: a physically meaningful equation must have the same units on both sides as they are connected by an identity sign. If you have [N] on one side there must be [N] on the other side, there cannot be [N] on one side and [Nm] on the other side - but this is exactly what is happening here. Frequency is "characterised by the number of occurrences of a repeating event per unit time." which was "Cycles per second" until 1960, when the SI unit was changed to "Hertz" in "per second" leaving out the "cycle".... that is as unserious as leaving out the "meter" from the unit of velocity which is "meter per second" not "per second".

So nowadays you get a dimensional analysis of Joule = Joule-second times per second $J=Js/s$ which reduces correctly to $J=J$, but that is flawed because frequency is "cycles per second" and then you get Joule = Joule-second times cycles per second $J=Js\cdot\text{cps}$ and that reduces to $J=J\cdot\text{cycles}$... which is a dimensionally unbalanced equation. The definition of frequency is still "the number of occurrences in time" so you have to name the occurrence and put it in, you can't leave it out because a "cycle" is physical information just like a "meter" - and therefore this equation unphysical.

So the question arises: why did this broken equation work for Planck in his paper on blackbody radiation? and how did Planck arrive at the value for [h]?

Normal Spectrum

Max Planck

Annalen der Physik vol.4, p.553 ff (1901)

The recent spectral measurements made by O. Lummer and E. Pringsheim[1], and even more notable those by H. Rubens and F. Kurlbaum[2], which together confirmed an earlier result obtained by H. Beckmann,[3] show that the law of energy distribution in the normal spectrum, first derived by W. Wien from molecular-kinetic considerations and later by me from the theory of electromagnetic radiation, is not valid generally.

In any case the theory requires a correction, and I shall attempt in the following to accomplish this on the basis of the theory of electromagnetic radiation which I developed. For this purpose it will be necessary first to find in the set of conditions leading to Wien's energy distribution law that term which can be

on page 6 of this paper we find the debut of $\epsilon = h\nu$

$$dU = n u \nu^{-1} d\nu$$

and integrated:

$$S = f\left(\frac{U}{\nu}\right) \tag{10}$$

that is, the entropy of a resonator vibrating in an arbitrary diathermic medium depends only on the variable U/ν , containing besides this only universal constants. This is the simplest form of Wien's displacement law known to me.

§10. If we apply Wien's displacement law in the latter form to equation (6) for the entropy S , we then find that the energy element ϵ must be proportional to the frequency ν , thus:

$$\epsilon = h\nu$$

and consequently:

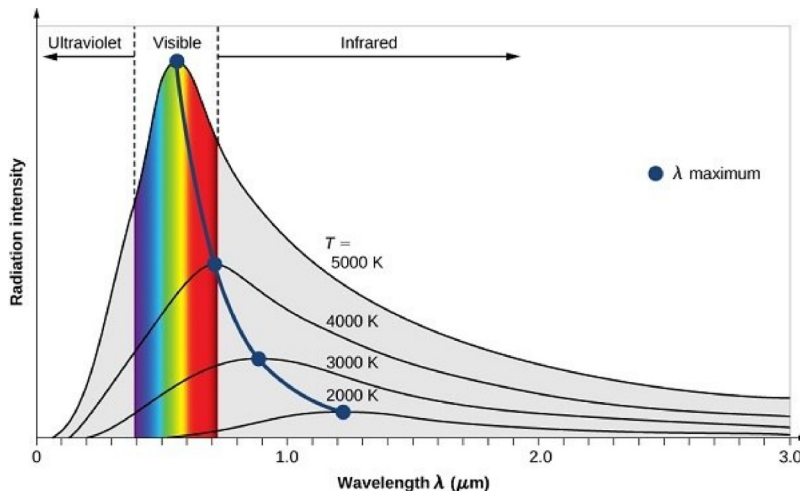
$$S = k \cdot \left\{ \left(1 + \frac{U}{h\nu}\right) \log \left(1 + \frac{U}{h\nu}\right) - \left(\frac{U}{h\nu}\right) \log \left(\frac{U}{h\nu}\right) \right\}$$

here h and k are universal constants.

By substitution into equation (9) one obtains:

$$\frac{1}{\theta} = \frac{k}{h\nu} \log \left(1 + \frac{h\nu}{U}\right)$$

Let's not forget that this is an attempt to find a mathematical solution that fits experimental data, and it so happens that these data sampled for the black body radiation curve was one seconds worth of data.



In chapter 3 about numerical values in this paper we see how that plays out:



5 Numerical values

§11. The values of both universal constants h and k may be calculated rather precisely with the aid of available measurements. F. Kurlbaum,[14] designating the total energy radiating into air from 1 cm² of a black body at temperature t °C in 1 sec by S_t , found that:

$$S_{100} - S_0 = 0.0731 \text{ watt/cm}^2 = 7.31 \cdot 10^5 \text{ erg/cm}^2\text{sec}$$

From this one can obtain the energy density of the total radiation energy in air at the absolute temperature 1:

$$\frac{4 \cdot 7.31 \cdot 10^5}{3 \cdot 10^{10}(373^4 - 273^4)} = 7.061 \cdot 10^{-15} \text{ erg/cm}^3\text{deg}^4$$

On the other hand, according to equation (12) the energy density of the total radiant energy for $q = 1$ is:

$$u^* = \int_0^\infty u dv = \frac{8\pi h}{c^3} \int_0^\infty \frac{v^3 dv}{e^{hv/k} - 1}$$

We see here that samples are taken "in 1 second" and the unit is Watt, which is "Joule per second", which corresponds to the 1s time lock of "Cycles per second" of the quantum equation $E=hv$. The old unit for Joule was [erg] and so we see here (erg/s)/cm²

Farther down we see the numerical value of both [h] and [k] arrived at:

On the other hand, it follows from equation (13), when one sets the derivative of E with respect to q equal to zero, thereby finding $\lambda = \lambda_m$:

$$\left(1 - \frac{ch}{5k\lambda_m\theta}\right) \cdot e^{ch/k\lambda_m\theta} = 1$$

and from this transcendental equation:

$$\lambda_m q = \frac{ch}{4.9651k}$$

consequently:

$$\frac{h}{k} = \frac{4.9561 \cdot 0.294}{3 \cdot 10^{10}} = 4.866 \cdot 10^{11}$$

From this and from equation (14) the values for the universal constants become:

$$h = 6.55 \cdot 10^{-27} \text{ erg sec} \tag{15}$$

$$k = 1.346 \cdot 10^{-16} \text{ erg / deg} \tag{16}$$

These are the same number that I indicated in my earlier communication.

... in the units we see that [h] has still that unit of time lingering around, whereas [k] is energy content per degree of temperature.

It was Einstein who thought he could conclude from this paper of Planck that nature has a discontinuous, point like structure, that light is quantized. He writes this idea up in this paper from 1905

Doc. 14
 ON A HEURISTIC POINT OF VIEW CONCERNING THE PRODUCTION
 AND TRANSFORMATION OF LIGHT
 by A. Einstein
 [Annalen der Physik 17 (1905): 132-148]

There exists a profound formal difference between the theoretical conceptions physicists have formed about gases and other ponderable bodies, and Maxwell's theory of electromagnetic processes in so-called empty space. While we conceive of the state of a body as being completely determined by the positions and velocities of a very large but nevertheless finite number of atoms and electrons, we use continuous spatial functions to determine the electromagnetic state of a space, so that a finite number of quantities cannot be considered as sufficient for the complete description of the

in this paper he introduces Planck's [hv] in disguise:

The usual conception, that the energy of light is continuously distributed over the space through which it travels, meets with especially great difficulties when one attempts to explain the photoelectric phenomena; these difficulties are presented in a pioneering work by Mr. Lenard.^[1]

According to the conception that the exciting light consists of energy quanta of energy $(R/N)\beta\nu$, the production of cathode rays by light can be conceived in the following way. The body's surface layer is penetrated by

[P] Lenard, *Ann. d. Phys.* 8 (1902): 169 and 170.

....here $[R/N]$ is the Universal Gas Constant over Avogadro's Number, which is equal to Boltzmann's Constant $[k]$ and $[\beta]$ is a number given at 4.866×10^{-11} which in Planck's Paper is given as the quotient of Planck's Constant over Boltzmann's Constant $h/k = 4.866 \times 10^{-11}$, so we end up with $(h/k)\nu$ and that is Planck's $[E = h\nu]$. There is no scientific reason whatsoever to disguise $[h\nu]$ in this way other than to hide something or come across as profound and original, when in fact it is all "borrowed".

.... farther down in the paper we find the photoelectric effect described:

100

HEURISTIC VIEW OF LIGHT

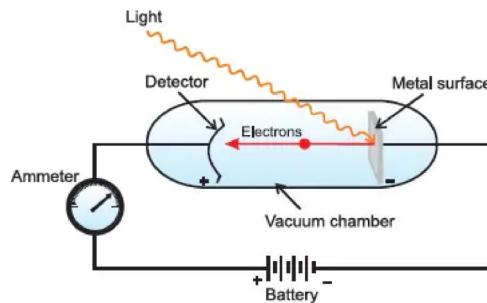
energy quanta whose energy is converted at least partially to kinetic energy of electrons. The simplest possibility is that a light quantum transfers its entire energy to a single electron; we will assume that this can occur.

However, we will not exclude the possibility that the electrons absorb only a part of the energy of the light quanta. An electron provided with kinetic energy in the interior of the body will have lost a part of its kinetic energy

[38] by the time it reaches the surface. In addition, it will have to be assumed that in leaving the body, each electron has to do some work P (characteristic for the body). The greatest perpendicular velocity on leaving the body will be that of electrons located directly on the surface and excited perpendicular to it. The kinetic energy of such electrons is


$$\frac{h}{N}\beta\nu - P.$$

here we see Planck's $[h\nu]$ minus $[P]$ which represents the work done by an electron to pass through matter while leaving the plate. For this there are practically only experimental data, although Owen Richardson tried to derive a formula for it.



But we know that $[h\nu]$ is a one second formula, that the energy of the photon is defined as one second's worth of energy. Yet we also know that it doesn't take one second's worth of time for the photoelectric effect to occur.... and the idea of quantum breaks down a second time.

But what is $[h]$ actually? is it really Joule-second?

Color	Wavelength	Frequency	Photon energy
violet	380–450 nm	668–789 THz	2.75–3.26 eV
blue	450–495 nm	606–668 THz	2.50–2.75 eV
green	495–570 nm	526–606 THz	2.17–2.50 eV
yellow	570–590 nm	508–526 THz	2.10–2.17 eV
orange	590–620 nm	484–508 THz	2.00–2.10 eV
red	620–750 nm	400–484 THz	1.65–2.00 eV

If we take a red light photon with 2eV energy we find it has 3.2×10^{-19} Joules at 484 Terahertz oscillations or cycles. That makes $3.2 \times 10^{-19} \text{ Joules} / 4.84 \times 10^{14}$ per cycle and that is 6.626×10^{-34} the numerical value of [h] except that here we get it in Joules-cycle and not in Joule-seconds

values

Constant	SI units	Units with eV
h	$6.626\ 070\ 15 \times 10^{-34} \text{ J} \cdot \text{Hz}^{-1}$ ^[2]	$4.135\ 667\ 696 \dots \times 10^{-15} \text{ eV} \cdot \text{Hz}^{-1}$ ^[5]
\hbar	$1.054\ 571\ 817 \dots \times 10^{-34} \text{ J} \cdot \text{s}$ ^[6]	$6.582\ 119\ 569 \dots \times 10^{-16} \text{ eV} \cdot \text{s}$ ^[7]
hc	$1.986\ 445\ 86 \dots \times 10^{-25} \text{ J} \cdot \text{m}$	$1.239\ 841\ 98 \dots \text{ eV} \cdot \mu\text{m}$
$\hbar c$	$3.161\ 526\ 77 \dots \times 10^{-26} \text{ J} \cdot \text{m}$	$0.197\ 326\ 9804 \dots \text{ eV} \cdot \mu\text{m}$

And here we see that physics is a language problem, not a math problem: if we say [h] is the quantum of energy of one EM oscillation the entire physics changes: now the wave and this mysterious particle become one and the same.

Let's see that play out in the equation:

$$E = h\nu$$

$$E = 6.626 \times 10^{-34} \text{ J} \cdot \text{Hz}$$

$$E = 6.626 \times 10^{-34} \text{ J} \cdot \text{Cycle/s}$$

$$E = 6.626 \times 10^{-34} \text{ J} \cdot \text{Cycles}$$

... and that means Planck's constant represents the energy of one EM cycle and NOT one particle, wave/particle duality is GONE! there is only the EM wave left.

This brings up another problem with photons, and that is the idea of high energy photons and low energy photons: When Planck's constant represents the energy of one EM oscillation, then there is no such thing as a high energy oscillation or low energy oscillation, [h] represents the constant Energy value of such an oscillation. So an oscillation of ultraviolet light has the same energy as an oscillation of infrared light, but the ultraviolet light hurts you and the infrared heals you. The answer is Capacitance: ultraviolet light has a higher Capacitance than infrared light, it has a higher counter spatial footprint and thus a "harder punch" per same amount of energy.

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Ali Abdulla · Mar 5

I think this is a waste of time discussion of what is well established empirically, although it's basic principles are weird. Many technological advances are due to quantum physics contributions.



2 Reply



Michael Brenner · Mar 5

That's an illusion, a delusion, nothing can be empirically established on the basis of irrationality and mathematical fraud.



Ali Abdulla I don't expect an engineer can comprehend quantum physics role in...



Davide Tanner Taini · Mar 11

Are you somewhat connected to Ken wheeler or what ?



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About the Author



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